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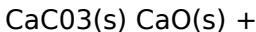
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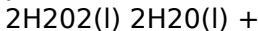
A sample of calcium carbonate is heated.



$\text{CO}_2(\text{g})$ Sulfur dioxide gas is bubbled through water.

$\text{SO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{SO}_3(\text{aq})$
Solid potassium oxide is added to a container of carbon dioxide gas.

$\text{K}_2\text{O}(\text{s}) + \text{CO}_2(\text{g}) \rightarrow \text{K}_2\text{CO}_3(\text{s})$
(s) Liquid hydrogen peroxide is warmed.



$\text{O}_2(\text{g})$ Solid lithium oxide is added to

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water.

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oxidation number of the underlined element: KMnO_4 . Since K is an alkali metal, its charge must be $1+$. Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$. If $1+$ and $8-$ are added together, we get $7-$. In order for the compound to be neutral, the Mn must be $7+$.

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