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## **The Periodic Table and Periodic LawThe Periodic Table and ...**

360 Chemistry: Matter and Change • Chapter 18 Solutions Manual CHAPTER 18 SOLUTIONS MANUAL Section 18.3 Hydrogen Ions and pH pages 650–658 Practice Problems pages 651–657 22. The concentration of either the  $H^+$  ion or the  $OH^-$  ion is given for four aqueous solutions at 298 K. For each solution, calculate  $[H^+]$  or  $[OH^-]$ . State whether the solution is acidic, basic, or neutral. a.  $[H^+] = 5.1 \times 10^{-3} M$  b.  $[OH^-] = 1.0 \times 10^{-5} M$

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